



## Abstract of Lectures on Chemistry (Paperback)

By William Russell Jones

Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1899 Excerpt: .Preparation:  $\text{Pb}(\text{NO}_3)_2 + 2\text{KI} = \text{PbI}_2 + 2\text{KNO}_3$ . Properties. Heavy, yellow powder, insoluble in water, soluble. in hot solution of ammonium chloride. Lead Sulphate,  $\text{PbSO}_4$ . Preparation. Made by precipitating lead nitrate with sodium sulphate:  $\text{Pb}(\text{NO}_3)_2 + \text{Na}_2\text{SO}_4 = 2\text{NaNO}_3 + \text{PbSO}_4$ . Properties. White, insoluble powder. Poisonous Properties of Lead Salts. Acute and chronic. Tests. 1. Hydrogen sulphide, or ammonium sulphides gives black precipitate, insoluble in dilute hydrochloric acid or ammonium sulphide. 2. Sulphates give a white precipitate. 3. Potassium iodide gives a yellow precipitate, insoluble in hot solution of ammonium chloride. 4. Potassium chromate gives a yellow precipitate. 5. Caustic potash gives a white precipitate, soluble in excess--plumbates,  $\text{K}_2\text{PbO}_3$ . 6. Ammonia water gives a white precipitate, insoluble in excess. 7. Dilute hydrochloric acid in cold, strong solutions of lead salts, gives white, granular precipitate, soluble in boiling water. COPPER. Symbol, Cu. Atomic Weight, 63.2 Valence, II....



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